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The first system involves a starch-iodine complex to demonstrate the effect of temperature on equilibrium. The formation of the complex is an exothermic reaction and results in a deep purple/blue-black color. Iodine (aq) + starch (aq)  $\rightleftharpoons$  starch-iodine complex (aq)  $\Delta H < 0$ . colorless.

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$q_{\text{reaction}} = - (4.18 \text{ J / g}\cdot\text{C} \times m_{\text{water}} \times \Delta t + C \times \Delta t)$   
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The lower the temperature, the slower the average speed of the molecules. When heat moves into a material it sometimes causes a change in temperature. Adding heat to a collection of molecules will cause their random motion to increase—the molecules will move faster on average than they did before heat was added.

Lab: Calorimetry - Chemistry

The water can be driven off by heating in an oven, leading to a measurable decrease in weight, which will gradually be regained if the fibre is returned to a 'normal' atmosphere. This effect is

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