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CHAPTER . 12 . REVIEW . Solutions . SHORT ANSWER Answer the following questions in the space provided. 1. The following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas in the liquid.

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Example 12.1.1. The solution in Figure 12.1.1 contains 10.0 g of cobalt(II) chloride dihydrate, $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}$, in enough ethanol to make exactly 500 mL of solution. What is the molar concentration of $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}$? Given: mass of solute and volume of solution Asked for: concentration (M) Strategy: To find the number of moles of $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}$, divide the mass of the compound by its

molar ...

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Solutions for Section 12.3 Solutions for Section 12.4 Solutions for Section 12.5 Solutions for Section 12.2 Exercise 12.2.1 SQL dates require 10 bytes, and SQL times require 8 bytes if there is no decimal point (See Section 12.1.3). a) The bytes required by each of the fields is $15 + 2 + 10 + 8 = 35$. b) Round each of the four field lengths up ...

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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

Chapter 12.2: Stoichiometry of Reactions in Solution ...

Example 12.1.1. The solution in Figure 12.1.1 contains 10.0 g of cobalt(II) chloride dihydrate, $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}$, in enough ethanol to make exactly 500 mL of solution. What is the molar concentration of $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}$? Given: mass of solute and volume of solution Asked for: concentration (M) Strategy: To find the number of moles of $\text{CoCl}_2 \cdot 2\text{H}_2\text{O}$, divide the mass of the compound by its molar ...

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