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Chapter 6 Test B The Periodic Table Answers

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.

Chapter 6 Practice Test Answers (1) - Chemistry Chapter 6 ...

Chapter Test B The Periodic

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Chapter Test B Chapter: The Periodic Law PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. In his periodic table, Mendeleev did not list all of the elements in order of increasing atomic mass because he wanted to group together elements with similar a. properties. b.

Chapter 6 The Periodic Table 143 Chapter Test A A. Matching Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line. B. Multiple Choice Choose the best answer and write its letter on the line. ____ 11. In the periodic table, there is a periodic pattern in the physical and

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Assessment Chapter Test B

Holt McDougal Modern Chemistry Chapter Test Assessment Chapter Test B Teacher Notes and Answers 5 The Periodic Law TEST B 1. a 2. c 3. d 4. d 5. a 6. a 7. c 8. a 9. lanthanides 10. 2 11. fourth 12. transition elements 13. 32 14. valence electrons 15. electron affinity 16. electronegativity 17. ionization energy 18. 3 s2 3p4 19. atomic radius 20. ion 21.

Chapter 6 Test B The Periodic Table Answers

3. The periodic law states that a. no two electrons with the same spin can be found in the same place in an atom. b. the physical and chemical properties of the elements are functions of their atomic number. c. electrons exhibit properties of both waves and particles. d. the chemical properties of elements can be grouped according to periodicity.

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Chapter 6 The Periodic Table Test B Answer Key | Elcho Table

Modern Chemistry 33 Chapter Test Name Class Date Chapter Test B, continued 15. The energy state of an atom is called ... The Periodic Law, pp. 36-45 TEST A 1. b 2. d 3. b 4. 5. d 6. a 7. b 8. 9. b 10. a 11. c 12. a 13. c 14. d 15. c 16. b 17. d 18. a 19. d 20. c 21. d 22. a 23. a 24. d 25. b TEST B

Assessment Chapter Test B - Ed W. Clark High School

____ 5. The Periodic Table of the Elements can be used by scientists A. to find out the main uses of each element. B. to predict how atoms of different elements will combine. C. to identify all of an element's physical and chemical properties. D. to determine the differences between ionic and covalent bonding.

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The part of an atom labeled B is a(an) a. proton. b. neutron. c. electron. ____ 22. Applying Concepts Which part of an atom has the least mass? a. A b. B c. C ____ 23. Applying Concepts Which parts of an atom make up nearly all the mass of an atom? a. A C b. A B c. B C - A B - C - - - - + - + + + + + Elements and the Periodic Table

Elements and the Periodic Table

Chapter Test A, continued ____ 3. The periodic law states that a. no two electrons with the same spin can be found in the same place in an atom. b. the physical and chemical properties of the elements are functions of their atomic number. c. wave patterns repeat at regular intervals.

Assessment Chapter Test A

Chapter 6 Practice Test Answers (1) - Chemistry Chapter 6 Test Matching Match each item with the correct statement below a electronegativity f b Chapter 6 Practice Test Answers (1) - Chemistry Chapter 6...

Chapter 6 Practice Test Answers (1) - Chemistry Chapter 6 ...

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c) atomic mass d) location in the periodic table c 2) The _____ model of an atom is a ball of positive charge with negatively charged electrons embedded in it.

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b. a neutron (Q an electron d. an ion 3. Across a period in the periodic table, ionization energy generally a. decreases. b. decreases, then increases. c increases. d. remains constant. 4. Down a group in

the periodic table, the change in ionization energy is due to @increasing electron shielding. b. decreasing charge of the nucleus.

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Chapter 6 Practice Test - Name Date Class THE PERIODIC... When elements are arranged in order of increasing atomic number, there is a periodic repetition of their physical and chemical properties. Column B a. electronegativity b. groups c. atomic radius d. ionization energy e. periodic law f. alkali metals g. halogens h. noble gases i. anion j.

Chapter 6 Practice Test - Name Date Class THE PERIODIC ...

Chapter 6 The Periodic Table 143 Chapter Test A A. Matching Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line. B. Multiple Choice Choose the best answer and write its letter on the line. ____ 11. In the periodic table, there is a periodic pattern in the physical and

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CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.

5 The Periodic Law

Mendeleev's table was called periodic because the properties of the elements. a. showed no pattern. b. occurred at repeated intervals called periods. c. occurred at regular time intervals called periods. d. were identical. ____ 10.

Chemistry Chapter 5 Exam

DIRECTIONS: Write on the line at the right of each statement the letter preceding the word or expression that best completes the statement. 1. Mendeleev did not always list elements in his periodic table in order of increasing atomic

The Periodic Law Chapter Test

Modern Chemistry 36 Chapter Test. Chapter: The Periodic Law. Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each. statement or best answers each question.

Assessment Chapter Test A

Chapter 7 The Periodic Table and Physical Properties LEVELED ASSESSMENT Chapter Review Chapter Tests BLTest A (Below Level) Test B (On Level) OL Test C (Advanced Learner) AL LABS For leveled labs, use the CD-ROM. Lab worksheets from Student Edition Labs MiniLab Lab: Version A (Below Level) BL Lab: Version B (On Level) OL (Advanced Learner) AL

Holt McDougal Modern Chemistry Chapter Test Assessment Chapter Test B Teacher Notes and Answers 5 The Periodic Law TEST B 1. a 2. c 3. d 4. d 5. a 6. a 7. c 8. a 9. lanthanides 10. 2 11. fourth 12. transition elements 13. 32 14. valence electrons 15. electron affinity 16. electronegativity 17. ionization energy 18. 3 s² 3p⁴ 19. atomic radius 20. ion 21.

Chapter 6 Practice Test Answers (1) - Chemistry Chapter 6 Test Matching Match each item with the correct statement below a electronegativity f b Chapter 6 Practice Test Answers (1) - Chemistry Chapter 6...

c) atomic mass d) location in the periodic table c 2) The _____ model of an atom is a ball of positive charge with negatively charged electrons embedded in it.

3. The periodic law states that a. no two electrons with the same spin can be found in the same place in an atom. b. the physical and chemical properties of the elements are functions of their atomic number. c. electrons exhibit properties of both waves and particles. d. the chemical properties of elements can be grouped according to periodicity.

DIRECTIONS: Write on the line at the right of each statement the letter preceding the word or expression that best completes the statement. 1. Mendeleev did not always list elements in his periodic table in order of increasing atomic

_____ 5. The Periodic Table of the Elements can be used by scientists A. to find out the main uses of each element. B. to predict how atoms of different elements will combine. C. to identify all of an element's physical and chemical properties. D. to determine the differences between ionic and covalent bonding.

Chapter 6 Practice Test - Name Date Class THE PERIODIC ...

Chapter 6 Practice Test - Name Date Class THE PERIODIC... When elements are arranged in order of increasing atomic number, there is a periodic repetition of their physical and chemical properties. Column B a. electronegativity b. groups c. atomic radius d. ionization energy e. periodic law f. alkali metals g. halogens h. noble gases i. anion j.

Assessment Chapter Test B

The part of an atom labeled B is a(an) a. proton. b. neutron. c. electron. ____ 22. Applying Concepts

Which part of an atom has the least mass? a. A b. B c. C ____ 23. Applying Concepts Which parts of an atom make up nearly all the mass of an atom? a. A C b. A B c. B C - A B - C - - - - + - + + + + + + Elements and the Periodic Table

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5 The Periodic Law

Elements and the Periodic Table

Modern Chemistry 33 Chapter Test Name Class Date Chapter Test B, continued 15. The energy state of an atom is called ... The Periodic Law, pp. 36-45 TEST A 1. b 2. d 3. b4. 5. d 6. a 7. b8. 9. b 10. a 11. c 12. a 13. c 14. d 15. c 16. b 17. d 18. a 19. d 20. c 21. d 22. a 23. a 24. d 25. b TEST B

Modern Chemistry 36 Chapter Test. Chapter: The Periodic Law. Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each. statement or best answers each question.

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b. a neutron (Q an electron d. an ion 3. Across a period in the periodic table, ionization energy generally a. decreases. b. decreases, then increases. c increases. d. remains constant. 4. Down a group in the periodic table, the change in ionization energy is due to @increasing electron shielding. b. decreasing charge of the nucleus.

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The Periodic Law ChapterTest

Chapter Test A, continued ____ 3. The periodic law states that a. no two electrons with the same spin can be found in the same place in an atom. b. the physical and chemical properties of the elements are functions of their atomic number. c. wave patterns repeat at regular intervals.

Assessment Chapter Test A

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