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Stoichiometry is the quantitative study of the relative amounts of reactants and products in chemical reactions; gas stoichiometry involves chemical reactions that produce gases. Stoichiometry is based on the law of conservation of mass, meaning that the mass of the reactants must be equal to the mass of the products.

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The stoichiometry of the reaction dictates that the number of moles CaCO₃ decomposed equals the number of moles CO₂ produced. Use the ideal-gas equation to convert moles of CO₂ to a volume. $V = nRT/P = (0.00150 \text{ mol}) (0.08206 \text{ L} \cdot \text{atm/mol} \cdot \text{K}) (273.15 \text{ K}) = 0.0336 \text{ L}$ or 33.6 mL.

The gas law stoichiometry is based on the number of moles of the substance. A chemical equation is said to be stoichiometrically balanced if the number of moles of the reactants and the products ...

Chemistry: Gas Laws and Stoichiometry. STUDY. PLAY. Kinetic Theory. a. The particles in a gas are considered to be small, hard spheres with an insignificant volume b. The motion of the particles in a gas is rapid, constant, and random c. All collisions between particles in a gas are perfectly elastic d. Gas particles travel in straight-line paths

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With an understanding of the ideal gas laws, it is now possible to apply these principles to chemical stoichiometry problems. For example, zinc metal and hydrochloric acid (hydrogen chloride dissolved in water) react to form zinc (II) chloride and hydrogen gas according to the equation shown below: $2 \text{HCl (aq)} + \text{Zn (s)} \rightarrow \text{ZnCl}_2 \text{ (aq)} + \text{H}_2 \text{ (g)}$

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When using stoichiometry with gas laws, you must remember: a) to use the ideal gas law . b) to use the combined gas law . c) all conditions are at STP . d) to convert to kilopascals. Ideal Gas ...

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10.5: Stoichiometry and the Ideal Gas Law - Chemistry ...

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